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**PHYSICS  
STANDARD LEVEL  
PAPER 2**

Thursday 10 May 2012 (afternoon)

1 hour 15 minutes

Candidate session number

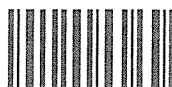
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Examination code

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**INSTRUCTIONS TO CANDIDATES**

- Write your session number in the boxes above.
- Do not open this examination paper until instructed to do so.
- Section A: answer all questions.
- Section B: answer one question.
- Write your answers in the boxes provided.
- A calculator is required for this paper.
- A clean copy of the **Physics Data Booklet** is required for this paper.
- The maximum mark for this examination paper is [50 marks].

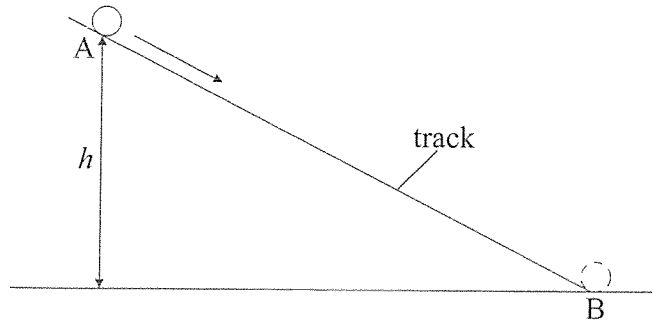


SECTION A

Answer *all* questions. Write your answers in the boxes provided.

A1. Data analysis question.

A small sphere rolls down a track of constant length AB. The sphere is released from rest at A. The time  $t$  that the sphere takes to roll from A to B is measured for different values of height  $h$ .

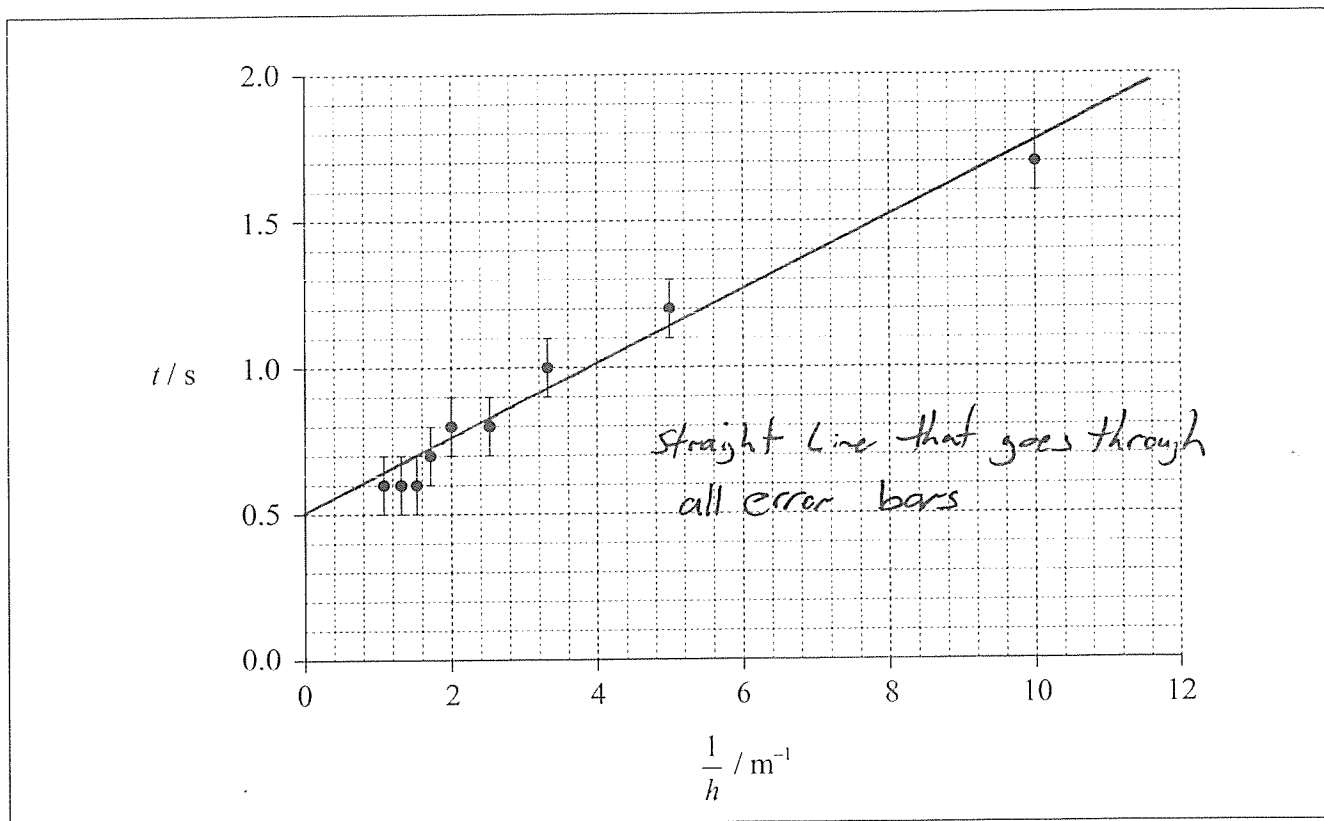


*(This question continues on the following page)*



(Question A1 continued)

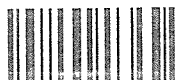
A student suggests that  $t$  is proportional to  $\frac{1}{h}$ . To test this hypothesis a graph of  $t$  against  $\frac{1}{h}$  is plotted as shown on the axes below. The uncertainty in  $t$  is shown and the uncertainty in  $\frac{1}{h}$  is negligible.



- (a) (i) Draw the straight line that best fits the data. [1]
- (ii) State why the data do not support the hypothesis. [1]

Line does not go through the origin.....  
.....

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(Question A1 continued)

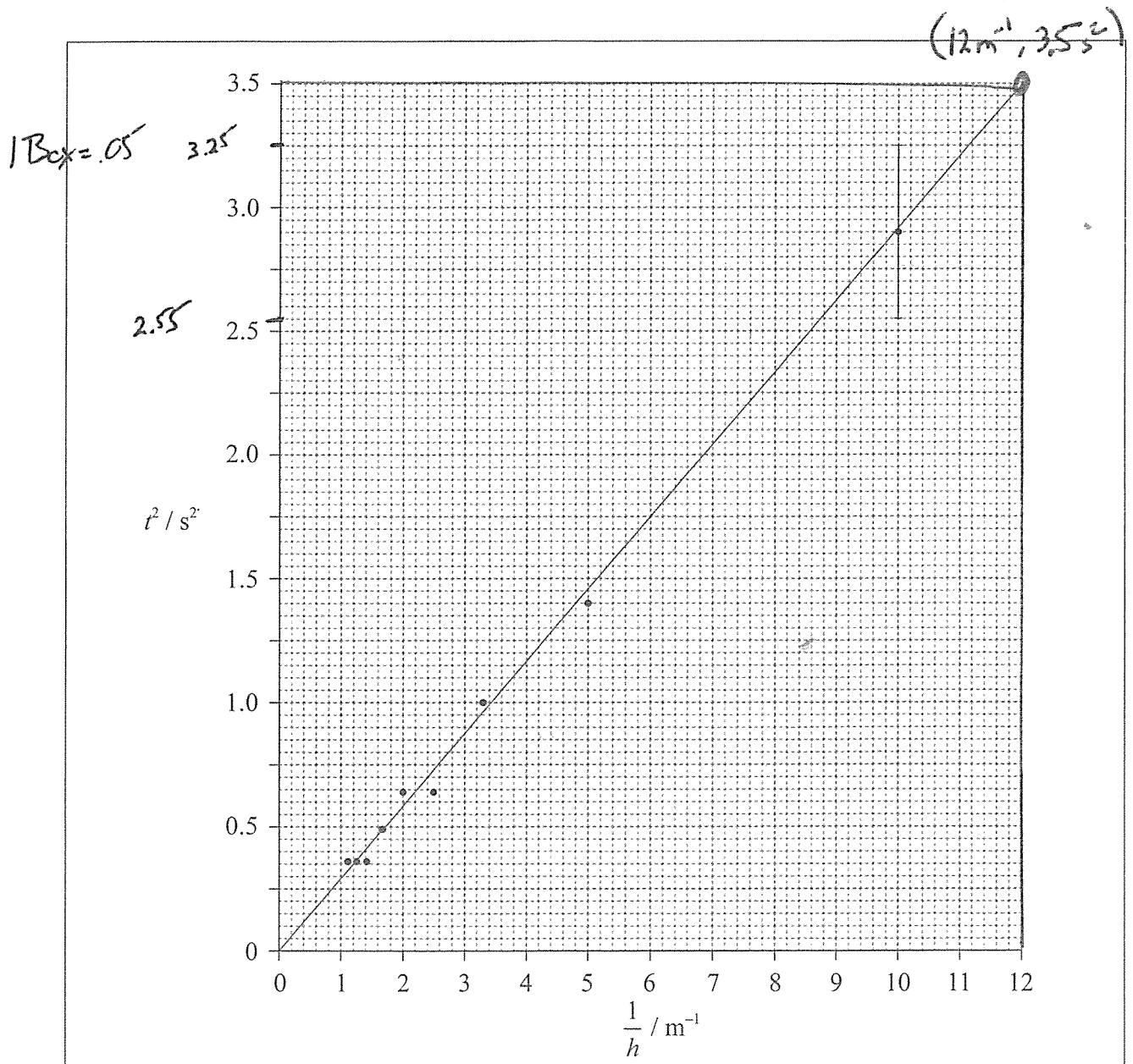
(b) Another student suggests that the relationship between  $t$  and  $h$  is of the form

$$t = k\sqrt{\frac{1}{h}}$$

where  $k$  is a constant.

To test whether or not the data support this relationship, a graph of  $t^2$  against  $\frac{1}{h}$  is plotted as shown below.

The best-fit line takes into account the uncertainties for all data points.



(This question continues on the following page)



(Question A1 continued)

The uncertainty in  $t^2$  for the data point where  $\frac{1}{h} = 10.0 \text{ m}^{-1}$  is shown as an error bar on the graph.

- (i) State the value of the uncertainty in  $t^2$  for  $\frac{1}{h} = 10.0 \text{ m}^{-1}$ . [1]

$$\frac{3.25 \text{ s}^2 - 2.55 \text{ s}^2}{2} = \frac{.70 \text{ s}^2}{2} = \pm .35 \text{ s}^2$$

- (ii) Calculate the uncertainty in  $t^2$  when  $t = 0.8 \pm 0.1 \text{ s}$ . Give your answer to an appropriate number of significant digits. [4]

$t^2 = (0.8 \pm 0.1 \text{ s}) \times (0.8 \pm 0.1 \text{ s})$  ... From Area & Perimeter Errors on E-Board

$$\frac{0.8 \text{ s}}{0.8 \text{ s}} \dots 1 \text{ SF} \quad \frac{\Delta Y}{Y} = \frac{.1}{0.8} + \frac{.1}{0.8}$$

$$.64 \text{ s}^2 = .65 \text{ s}^2$$

$$\frac{\Delta Y}{Y} = .25 \quad \Delta Y = .15 \text{ s}^2 = .25 \text{ s}^2$$

$$0.65 \text{ s}^2 \pm .25 \text{ s}^2 \quad \text{Uncertainty} = .25 \text{ s}^2$$

- (iii) Use the graph to determine the value of  $k$ . Do not calculate its uncertainty. [3]

Solve with equation or slope  $(t^2 = K \sqrt{\frac{1}{h}})$   $\sqrt{K^2} = \sqrt{.29}$

Plug in point from graph  $t^2 = K^2 \left(\frac{1}{h}\right)$   $K = .54$

Accept: .53 - .57  $\frac{3.5 \text{ s}^2}{12 \text{ m}^{-1}} = \frac{K^2 (12 \text{ m}^{-1})}{12 \text{ m}^{-1}}$

- (iv) State the unit of  $k$ . [1]

$$K^2 = \frac{\text{s}^2}{\text{m}^{-1}} \quad \sqrt{K^2} = \sqrt{\text{m} \text{ s}^2} \quad K = \text{m}^{1/2} \text{ s}$$



A2. This question is about the greenhouse effect.

The following data are available for use in this question:

Quantity	Symbol	Value
Power emitted by the Sun	$P$	$3.8 \times 10^{26} \text{ W}$
Distance from the Sun to the Earth	$d$	$1.5 \times 10^{11} \text{ m}$
Radius of the Earth	$r$	$6.4 \times 10^6 \text{ m}$
Albedo of the Earth's atmosphere	$\alpha$	0.31
Stefan-Boltzmann constant	$\sigma$	$5.7 \times 10^{-8} \text{ Wm}^{-2}\text{K}^{-4}$

(a) Explain why the power absorbed by the Earth is

*Need to explain each part  
of the equation for a point*

$$\frac{P}{4\pi d^2} \times (1 - \alpha) \times \pi r^2$$

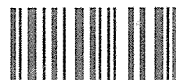
[3]

Intensity of Sun's radiation at the Earth's orbit =  $\frac{P}{4\pi d^2}$

Fraction absorbed by the earth  $(1 - \alpha)$

Surface area of the earth absorbing the radiation  $\pi r^2$

(This question continues on the following page)



(Question A2 continued)

- (b) The equation in (a) leads to the following expression which can be used to predict the Earth's average surface temperature  $T$ .

$$T = \sqrt[4]{\frac{(1-\alpha)P}{16\pi\sigma d^2}}$$

- (i) Calculate the predicted temperature of the Earth. [2]

$$T = \left( \frac{(1-0.31)(3.8 \times 10^{26} \text{ W})}{16\pi (5.7 \times 10^{-8} \text{ W m}^{-2} \text{ K}^{-4}) (1.5 \times 10^{11} \text{ m})^2} \right)^{1/4}$$

$$T = \left( \frac{2.6 \times 10^{26}}{6.4 \times 10^{16}} \right)^{1/4} = 252 \text{ K or } 250 \text{ K (25 F)}$$

- (ii) Explain why the actual average surface temperature of the Earth is in fact higher than the answer to (b)(i). [2]

Green House gases in the atmosphere trap some energy radiated by the earth and prevent it from leaving our earth.



A3. This question is about thermal energy transfer.

A hot piece of iron is placed into a container of cold water. After a time the iron and water reach thermal equilibrium. The heat capacity of the container is negligible.

(a) Define *specific heat capacity*.

[2]

$Q = mc\Delta T$  ...  $c =$  specific heat  
 = Amount of heat / energy required to raise the temp  
 ...  $1^\circ\text{C}$  per unit mass

(b) The following data are available.

Mass of water	= 0.35 kg
Mass of iron	= 0.58 kg
Specific heat capacity of water	= $4200 \text{ J kg}^{-1} \text{ K}^{-1}$
Initial temperature of water	= $20^\circ\text{C}$
Final temperature of water	= $44^\circ\text{C}$
Initial temperature of iron	= $180^\circ\text{C}$

(i) Determine the specific heat capacity of iron.

[3]

2 SF

$Q_{\text{iron}} = Q_{\text{water}}$  ...  $mc\Delta T$  ...  $mc\Delta T$   
 $0.58 \text{ kg} (c) (44^\circ\text{C} - 180^\circ\text{C}) = 0.35 \text{ kg} (4200 \text{ J kg}^{-1} \text{ K}^{-1}) (44^\circ\text{C} - 20^\circ\text{C})$   
 $c = 447 \text{ J kg}^{-1} \text{ K}^{-1} \approx 450 \text{ J kg}^{-1} \text{ K}^{-1}$

(ii) Explain why the value calculated in (b)(i) is likely to be different from the accepted value.

[2]

Not a closed system - energy would be given off  
 ... to the surroundings i.e. the container  
 Final temp of iron would be less, so the specific heat  
 of the iron would be greater.

